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0.35 «cm³ s⁻¹» Accept values in the range

0.32-0.42 «cm³ s⁻¹».

3 2. b ALTERNATIVE 1 colour Br₂ /reactant is coloured

«Br (aq) is not» ALTERNATIVE 2

conductivity

greater/increased

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concentration of ions in
products ALTERNATIVE
3 mass/pressure gas is
evolved/produced
ALTERNATIVE 4 pH

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3. CH₂, 2. CHO contains
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a (strong) peak at $m/z = 57$; CH_3COCH_3 .

does . not. contain a (strong) peak at $m/z = 57$; (corresponding to) loss of $\text{H} / \text{M} . \text{r} - \text{H} / \text{CH}_3 . \text{CH}_2 . \text{CO} +$; [2]

Penalize missing + once only in A1. (d)

$m/z = 71$: $\text{CH}_3\text{CH}_2\text{CH}_2\text{CO} + / \text{C}_3\text{H}_7\text{CO} + / \text{C}_4\text{H}_7\text{O} +$; Accept $\text{CH}_3\text{COCH}_2\text{CH}_2 +$.

$m/z = 43$: $\text{CH}_3 . \text{CH}_2 . \text{CH} \dots$

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Chemistry in industry
and technology 9. (a)

(i) + 8 3 / 2 2 3 + ;

Accept +2.7 but not
+3. +2 and +3 /

contains two (or more)
iron ions with different
oxidation states /

contains Fe²⁺ and

Fe³⁺; [2] Accept II and

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III oxidation number notation for oxidation states but not $2+$ and $3+$ unless ions are referred to explicitly.
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hl CHEMISTRY NOTES.

TOPIC 1:

QUANTITATIVE

CHEMISTRY. TOPIC 2:

ATOMIC STRUCTURE ...

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